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Holt ChemFile: Problem-Solving Workbook 58 Mole Concept Name Class Date Problem Solving continued CONVERTING NUMBER OF ATOMS OF AN ELEMENT TO MASS Sample Problem 4 uses the progression of steps 1→2→3 to convert from the mass of an element to the number of atoms. In order to calculate the mass from a given number of atoms, these steps will be reversed.

Skills Worksheet Problem Solving

Holt ChemFile: Problem-Solving Workbook 118 Limiting Reactants Name Class Date Problem Solving continued Only 0.0121 mol Cl₂ is available. For all of the KBr to react, 0.0136 mol Cl₂ is needed. Therefore, Cl₂ is the limiting reactant. EVALUATE Is the determination of limiting reactant reasonable? Yes; the mass of 2 mol KBr is 238 g and the mass of 1 mol Cl

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Holt ChemFile: Problem-Solving Workbook 99 Stoichiometry Name Class Date Problem Solving continued Sample Problem 1 Ammonia is made industrially by reacting nitrogen and hydrogen under pressure, at high temperature, and in the presence of a catalyst. The equation is N₂(g) + 3H₂(g)→2NH₃(g). If 4.0 mol of H₂ react, how many moles of NH₃ will be produced?

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